

# Additions and Corrections

## Determination of Standard Gibbs Energies of Transfer of Organic Anions across the Water/Nitrobenzene Interface

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Some of the published data in Table 3 of the paper cited above are not correct because contamination of the decamethylferrocene caused an erroneous electrode response, i.e., a shift of the voltammetric peak potentials. Careful re-examination of all systems using decamethylferrocene of highest purity led to the data given in Table 1 of this correction. The new  $\Delta G^{\circ}_{X^{-}(W \rightarrow NB)}$  data deviate from the erroneously reported ones especially in the case of the anions valerate, capronate, and oenanthate. For the other anions listed in Table 1, the deviations are marginal; however, we feel that the correct values must be reported. The technique used provides a standard deviation ( $n = 12$ ) of the free energies of  $0.16 \text{ kJ mol}^{-1}$ . As  $\log(P)$  values ( $P$  is the partition coefficient) are usually used for assessing the lipophilicity of compounds, it is worth mentioning that the new data of  $\Delta G^{\circ}_{X^{-}(W \rightarrow NB)}$  give only slightly different new  $\log(P)$  values. Re-examination of all other data reported in the previous publication proved that they are completely correct.

**Table 1. Peak Potentials of the Net SW Voltammetric Response, Standard Potential Differences at the W/NB Liquid Interface, and the Standard Gibbs Energies of the Transfer of the Anions of Aliphatic Monocarboxylic Acids<sup>a</sup>**

anion	$E_p/\text{mV vs Ag/AgCl}$	$\Delta_{W^{NB}}\varphi^{\circ}_{X^{-}}/\text{mV}$	$\Delta G^{\circ}_{X^{-}(W \rightarrow NB)}/\text{kJ mol}^{-1}$	
			new data	old data
formiate	58	317.10	30.60	28.07
acetate	52	311.92	30.10	29.14
propionate	29	289.94	27.98	27.30
butyrate	11	272.02	26.25	27.86
valerate	-31	231.09	22.30	26.82
capronate	-75	187.56	18.10	22.94
oenanthate	-115	147.15	14.20	17.88
caprylate	-125	130.98	12.64	13.78
pelargonate	-120	138.86	13.40	13.02
caprinate	-118	140.93	13.60	12.71

<sup>a</sup> The concentration of the anions in the aqueous phase was  $1 \text{ mol dm}^{-3}$ . All other conditions were the same as in Figure 1. The values of  $\Delta_{W^{NB}}\varphi^{\circ}_{X^{-}}$  and  $\Delta G^{\circ}_{X^{-}(W \rightarrow NB)}$  are estimated using the calibration line in Figure 2.

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