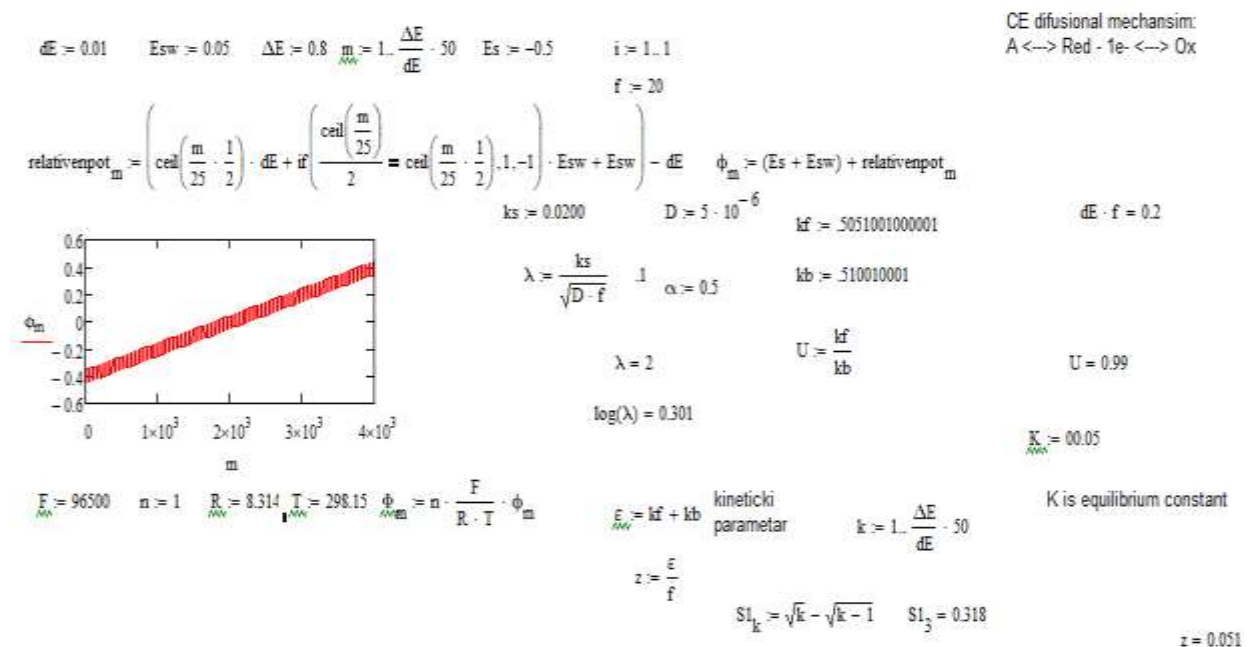


# Effect of Equilibrium Constant of Chemical Reaction to the features of square-wave voltammograms of a diffusional "CrevErev" mechanism featuring preceding homogeneous chemical step

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**Abstract:** The electrochemical diffusional mechanism featuring reversible homogeneous chemical reaction (diffusional "CrevErev" mechanism) is solved theoretically under conditions of square-wave voltammetry. In this file, a MATHCAD simulation protocol of this important electrochemical mechanism is presented, while highlighting the effect of equilibrium constant of the chemical step to the features of all current components of simulated Square-wave voltammograms. Method for the determination of equilibrium constant (Keq) is proposed. MATHCAD file provided is free and ready to use.

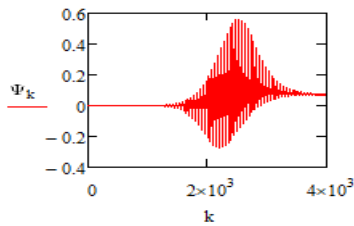


$$S_{kz} = \text{erfc} \left[ z \cdot \frac{k}{50} \right]^{0.5} - \text{erfc} \left[ z \cdot \frac{(k-1)}{50} \right]^{0.5}$$

$$\Psi_1 = \lambda \cdot e^{-\alpha \cdot \phi_1} \cdot \frac{K}{1+K} \cdot \left[ 1 + \lambda \cdot e^{-\alpha \cdot \phi_1} \cdot \frac{K}{(1+K) \cdot 50} - \frac{\lambda \cdot e^{-\alpha \cdot \phi_1} \cdot S_1}{(K+1) \cdot z} \cdot (1) + \frac{\lambda \cdot e^{-(1-\alpha) \cdot \phi_1}}{50} \right]^{-1}$$

$$\Psi_k := \frac{\frac{\lambda \cdot e^{\alpha \cdot \Phi_k} \cdot K}{1 + K} \cdot \left[ 1 - \frac{2}{\sqrt{50 \cdot \pi}} \cdot \sum_{j=1}^{k-1} (\Psi_j \cdot S_{1_{k-j+1}}) \right] - \gamma \cdot \left( \frac{1}{1 + K} \right) \cdot (-1) \cdot e^{\alpha \cdot \Phi_k} \cdot \sum_{j=1}^{k-1} (\Psi_j \cdot S_{k-j+1}) - \lambda \cdot \frac{2}{\sqrt{50 \cdot \pi}} \cdot e^{-\Phi_k \cdot (1-\alpha)} \cdot \sum_{j=1}^{k-1} (\Psi_j \cdot S_{1_{k-j+1}})}{\left( \frac{\lambda \cdot e^{\alpha \cdot \Phi_k} \cdot K}{1 + K} \cdot \frac{2 \cdot S_{1_1}}{\sqrt{50 \cdot \pi}} \right) + 1 + \gamma \cdot (-1) \cdot \left( \frac{1}{1 + K} \right) \cdot S_1 \cdot e^{\alpha \cdot \Phi_k} + \lambda \cdot \frac{2 \cdot S_{1_1}}{\sqrt{50 \cdot \pi}} \cdot e^{-\Phi_k \cdot (1-\alpha)}}$$

$$\Psi_1 = -$$



$$p := 1.. \frac{\Delta E}{dE} - 2$$

$$\Psi f_p := (\Psi)_{(p+1) \cdot 50}$$

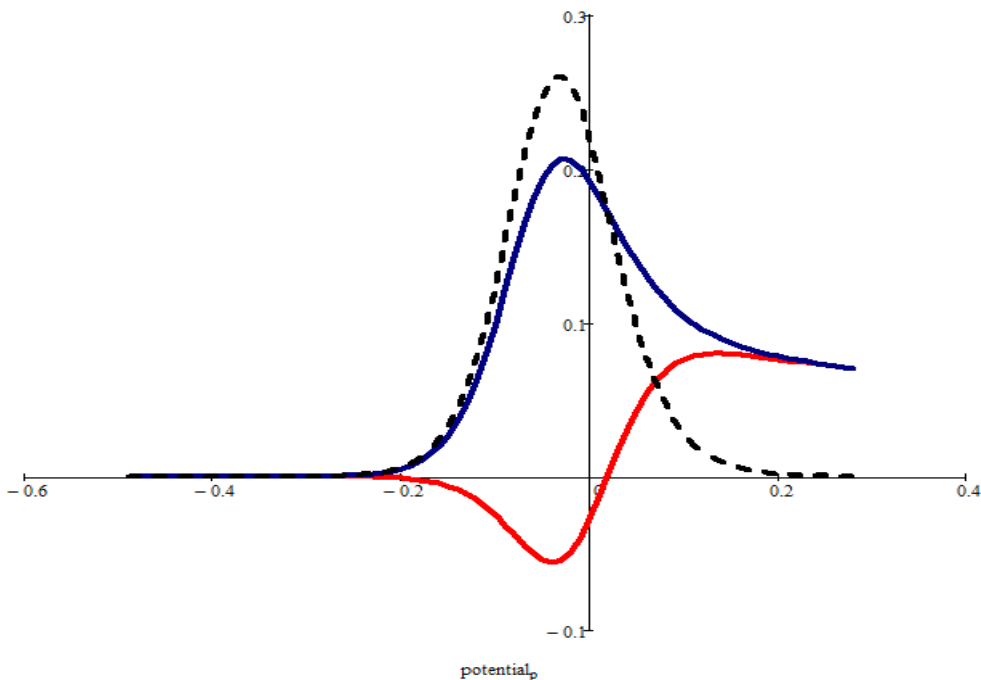
$$\Psi b_p := (\Psi)_{50 \cdot p + 25}$$

$$\Psi net_p := \Psi f_p - \Psi b_p$$

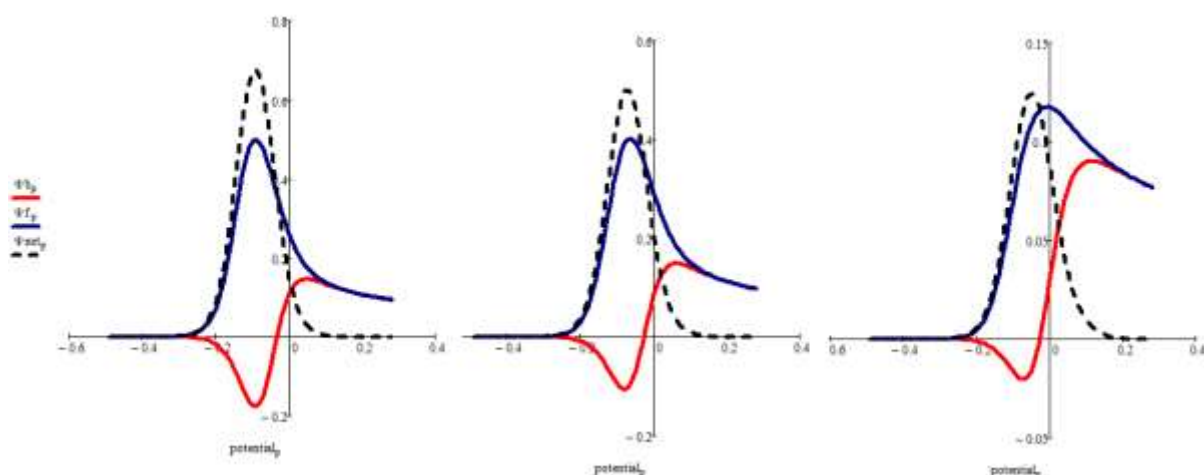
$$potential_p := E_s + (p) \cdot dE$$

$$\Psi_{k,i} =$$

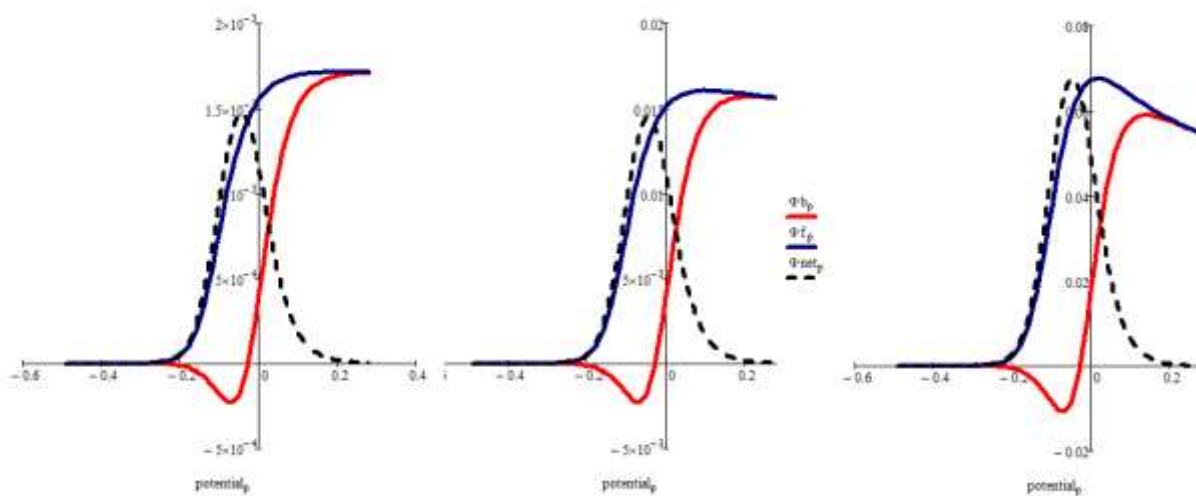
$\Psi b_p$   
 $\Psi f_p$   
 $\Psi net_p$



- $6.747 \times 10^{-}$
- $9.253 \times 10^{-}$
- $1.33 \times 10^{-}$
- $1.94 \times 10^{-}$
- $2.848 \times 10^{-}$
- $4.192 \times 10^{-}$
- $6.179 \times 10^{-}$
- $9.113 \times 10^{-}$
- $1.344 \times 10^{-}$
- $1.984 \times 10^{-}$
- $2.928 \times 10^{-}$
- $4.322 \times 10^{-}$
- $6.38 \times 10^{-}$
- $9.418 \times 10^{-}$
- $1.39 \times 10^{-}$
- $2.052 \times 10^{-}$
- $3.03 \times 10^{-}$
- $4.474 \times 10^{-}$
- $6.605 \times 10^{-}$
- $9.751 \times 10^{-}$
- $1.44 \times 10^{-}$
- $2.125 \times 10^{-}$
- $3.137 \times 10^{-}$
- $4.628 \times 10^{-}$
- $6.825 \times 10^{-}$
- $1.006 \times 10^{-}$
- $1.48 \times 10^{-}$



Effect of equilibrium constant of preceding chemical reaction the features of simulated square-wave voltammograms of a Cerev mechanism at high rates of chemical step. SW voltammograms are calculated for  $K_{eq} = 5.5$ ;  $0.5$  and  $0.05$  (from left to right)



Effect of equilibrium constant of preceding chemical reaction the features of simulated square-wave voltammograms of a Cerev mechanism at high rates of chemical step and low values of  $K_{eq}$ . SW voltammograms are calculated for  $K_{eq} = 0.00005$ ;  $0.0005$  and  $0.025$  (from left to right)

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