

Electrochemically Induced Surface Dimerization in SWV-MATHCAD

Rubin Gulaboski, Goce Delcev University Stip, MACEDONIA

$$E_{sl} := 0.3 \quad \Delta E := 0.8 \quad dE := 0.004 \quad E_{sw} := 0.085$$

$$n := 1 \quad F := 96500 \quad R := 8.314 \quad T := 298.15 \quad \alpha := 0.5$$

$$k_s := 2.5 \quad k_f := 0.0005 \quad c_{Ox} := 10.0 \quad r_s := 1$$

$$k_b := 0.0005 \quad c_s := \frac{c_{Ox}}{r_s}$$

$$f := 10$$

$$K_{eq} := 1.00$$

MATHCAD FILE

ELECTROCHEMICALLY INDUCED DIMERIZATION IN LIPOPHILIC REDOX PROTEINS
THEORETICAL INSIGHTS IN SQUARE-WAVE VOLTAMMETRY

$$j := 1.. \frac{\Delta E}{dE} \cdot 50$$

$$pot_j := E_{sl} + E_{sw} - \left[\left(\frac{ceil\left(\frac{j}{25}\right)}{25} \cdot \frac{1}{2} \right) dE + \text{if} \left(\frac{ceil\left(\frac{j}{25}\right)}{2} = \text{ceil}\left(\frac{j}{25}\right), 1, -1 \right) \cdot E_{sw} + E_{sw} \right] - dE$$

$$K_{et} := \frac{k_s}{f}$$

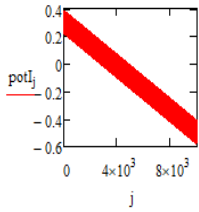
$$K_{chem} := \frac{(k_f + k_b)}{f}$$

$$M_j := e^{-K_{chem} \cdot \frac{j}{50}} - e^{-K_{chem} \cdot \frac{j+1}{50}}$$

Surface E-dimerization mechanism in SWV
 $Ox(ads) + 1e^- = Red(ads)$
 $2Red(ads) \rightleftharpoons D$

Meaning of symbols in the MATHCAD File

- k_s** is standard rate constant of electron transfer (s⁻¹)
- k_f** is forward rate constant of dimerization step
- k_b** is reverse rate constant of dimerization step
- f** is the SW frequency (Hz)
- E_{sw}** is SW amplitude (V)
- dE** is potential step (V)
- n** is number of electrons exchanged
- α** is electron transfer coefficient
- F** is Farady constant
- R** is universal gas constant
- T** is thermodynamic temperature
- c_{Ox}** is molar concentration of Ox
- r_s** is auxiliary constant
- c_s** is dimensionless concentration parameter
- K_{chem}** is chemical rate parameter of dimerization reaction
- K_{eq}** is equilibrium constant of dimerization reaction
- M_j** is numerical integration factor
- Φ_j** is dimensionless potential
- E_{sl}** is starting potential
- K_{et}** is dimensionless parameter related to electron transfer step
- Ψ** is symbol of the dimensionless current



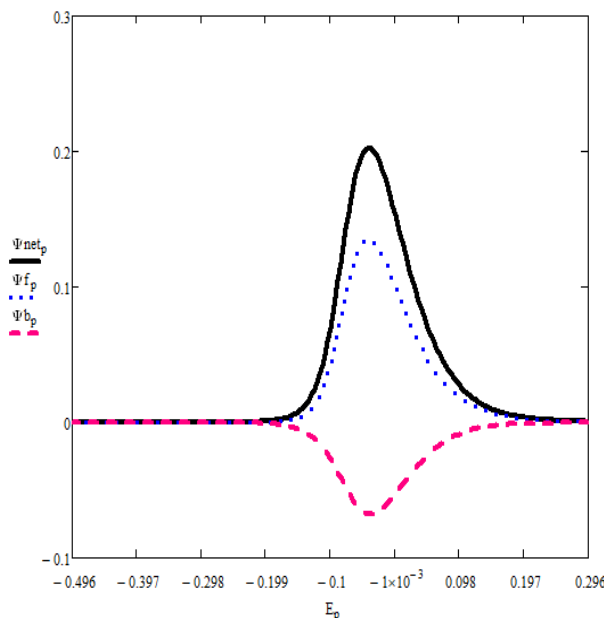
$$\Phi_j := n \frac{F}{1R \cdot T} \cdot pot_j$$

$$\Psi_{I_1} := \frac{K_{et} \cdot e^{-\alpha \cdot \Phi_{I_1}}}{1 + \frac{K_{et}}{50} \cdot e^{-\alpha \cdot \Phi_{I_1}} + \frac{1}{c_s} \cdot K_{et} \cdot K_{chem}^{-1} \cdot e^{-\Phi_{I_1} \cdot (1-\alpha)} \cdot M_1}$$

$$\Psi_j := \frac{\frac{K_{et}}{c_s} \cdot e^{-\alpha \cdot \Phi_j} - \frac{K_{et}}{50 \cdot c_s} \cdot e^{-\alpha \cdot \Phi_j} \sum_{i=1}^{j-1} \Psi_{I_i} - \left(\frac{1}{1}\right) \left[1 - \frac{K_{et}}{1} \cdot K_{chem}^{-1} \cdot e^{-\Phi_j \cdot (1-\alpha)} \cdot \sum_{i=1}^{j-1} (\Psi_{I_i} \cdot M_i) - \frac{1 \cdot K_{eq}}{1 + K_{eq}} \cdot K_{chem}^{-1} \cdot e^{-\Phi_j \cdot (1-\alpha)} \cdot \sum_{i=1}^{j-1} (\Psi_{I_i} \cdot M_i) \right]^2}{1 + \frac{K_{et}}{50 \cdot c_s} \cdot e^{-\alpha \cdot \Phi_j} + \frac{1}{1} \left[1 - \frac{K_{et}}{1} \cdot K_{chem}^{-1} \cdot e^{-\Phi_j \cdot (1-\alpha)} \cdot M_1 - \frac{1 \cdot K_{eq}}{1 + K_{eq}} \cdot K_{chem}^{-1} \cdot e^{-\Phi_j \cdot (1-\alpha)} \cdot M_1 \right]^2}$$

$$\Psi_j := \Psi_{I_j} \quad p := 1.. \left(\frac{\Delta E}{dE} \right) - 1$$

$$\Psi_{f_p} := \Psi_{50 \cdot p + 25} \quad \Psi_{b_p} := \Psi_{(p+1) \cdot 50} \quad \Psi_{net_p} := \Psi_{f_p} - \Psi_{b_p} \quad E_p := E_{sl} - p \cdot dE$$



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